

30/11/93

Internal Memo



To ~~Mr Campbell~~
Customer Projects
Executive Office - C&C

From ~~Mr [redacted]~~

Commercial & Consumer
Customer Affairs

Locked Bag 4960
Melbourne Vic 3100

Subject TIO AND COT

File

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Date 30 November, 1993

At today's Council Meeting the TIO reported on his involvement with the COT settlement processes - it was agreed that any financial contributions made by Telecom to the COT arbitration process was not a matter for Council but was a private matter between Telecom, AUSTEL and the TIO.

I hope you agree with this.



GROUP MANAGER - CUSTOMER AFFAIRS

FAXED
2/12/93

Don

seems OK to me.

when I spoke to Darwin I suggested that at least for the first group he develops a separate budget, clears with us, approves bills for payment and we pay direct.

when we have had experience with this can decide to continue for Group 2 and future

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001248

Telstra Corporation Limited
ACN 051 775 556

Journal

DATE: _____

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CLASS: _____

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The first part of the experiment was to determine the effect of temperature on the rate of reaction. The reaction was carried out at three different temperatures: 25°C, 35°C, and 45°C. The rate of reaction was measured by the time taken for the reaction to complete. The results showed that the rate of reaction increased as the temperature increased.



$$\frac{1}{t} \propto k$$



It was found that the rate of reaction was directly proportional to the temperature. This is because the rate constant, k , increases with temperature. The Arrhenius equation, $\ln k = \ln A - \frac{E_a}{RT}$, shows that $\ln k$ is a linear function of $1/T$. The slope of this line is $-E_a/R$, where E_a is the activation energy and R is the gas constant.

The activation energy, E_a , is the minimum energy required for a reaction to occur. It is a characteristic property of a reaction and is independent of the concentration of the reactants. The activation energy can be determined from the slope of the Arrhenius plot.

The rate of reaction was also found to be dependent on the concentration of the reactants. The rate of reaction was directly proportional to the concentration of the reactants. This is because the rate constant, k , is independent of the concentration of the reactants.

The rate of reaction was also found to be dependent on the surface area of the reactants. The rate of reaction was directly proportional to the surface area of the reactants. This is because the rate constant, k , is independent of the surface area of the reactants.

The rate of reaction was also found to be dependent on the presence of a catalyst. The rate of reaction was directly proportional to the concentration of the catalyst. This is because the rate constant, k , is independent of the concentration of the catalyst.

The rate of reaction was also found to be dependent on the pressure of the reactants. The rate of reaction was directly proportional to the pressure of the reactants. This is because the rate constant, k , is independent of the pressure of the reactants.